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According to the following reaction:

$$3H_2 + N_2 \rightarrow 2NH_3$$

What is the partial pressure of NH_3 if one starts with 33 atm of H_2 and an excess of N_2 and the reaction goes to completion? The volume and temperature of the container is the same after the reaction as before.

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What volume of SF_6 can be produced at STP if one starts the following reaction with 45.0 L of F_2 at STP?

$$S_8 + 24F_2 \rightarrow 8SF_6$$



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What volume of CO_2 can be produced at STP if one starts the following reaction with 39.0 L of O_2 at STP?

$$2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$$



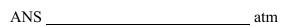
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What pressure of CO_2 can be produced in a constant volume container at constant temperature if one starts the following reaction with 23.0 atm of O_2 ?

$$2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$$



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What volume of CF_4 can be produced at STP if one starts the following reaction with 25.0 L of F_2 at STP?

$$2CO + 4F_2 \rightarrow 2CF_4 + O_2$$

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What pressure of CCl_4 can be produced in a constant volume container at constant temperature if one starts the following reaction with 13.0 atm of Cl_2 ?

$$C_2H_4 + 6Cl_2 \rightarrow 2CCl_4 + 4HCl$$

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Critical Question 10

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What volume of $H_2O(g)$ is produced when 27.0 L of O_2 reacts at constant pressure and temperature with an excess of H_2 according to the reaction:

$$2H_2 + O_2 \rightarrow 2H_2O$$
 ?



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What pressure of $H_2(g)$ is required to react in precise stoichiometry with 27.0 atm of O_2 in a constant volume and constant temperature? The reaction is:

$$2H_2 \ + \ O_2 \ \, {\color{red} \rightarrow} \ \, 2H_2O \; .$$

