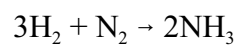


Critical Question 10

NAME _____

According to the following reaction:



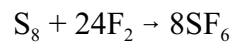
What is the partial pressure of NH_3 if one starts with 33 atm of H_2 and an excess of N_2 and the reaction goes to completion? The volume and temperature of the container is the same after the reaction as before.

ANS _____ atm

Critical Question 10

NAME _____

What volume of SF₆ can be produced at STP if one starts the following reaction with 45.0 L of F₂ at STP?

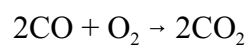


ANS _____ L

Critical Question 10

NAME _____

What volume of CO₂ can be produced at STP if one starts the following reaction with 39.0 L of O₂ at STP?

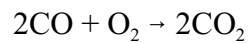


ANS _____ L

Critical Question 10

NAME _____

What pressure of CO₂ can be produced in a constant volume container at constant temperature if one starts the following reaction with 23.0 atm of O₂?

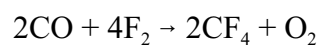


ANS _____ atm

Critical Question 10

NAME _____

What volume of CF_4 can be produced at STP if one starts the following reaction with 25.0 L of F_2 at STP?

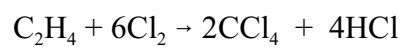


ANS _____ L

Critical Question 10

NAME _____

What pressure of CCl_4 can be produced in a constant volume container at constant temperature if one starts the following reaction with 13.0 atm of Cl_2 ?

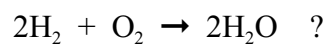


ANS _____ atm

Critical Question 10

NAME _____

What volume of $\text{H}_2\text{O}(\text{g})$ is produced when 27.0 L of O_2 reacts at constant pressure and temperature with an excess of H_2 according to the reaction:

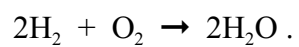


_____ L

Critical Question 10

NAME _____

What pressure of H₂(g) is required to react in precise stoichiometry with 27.0 atm of O₂ in a constant volume and constant temperature? The reaction is:



_____ atm