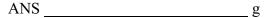
Critical Item 4 - created on June 12	, 2008
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Calculate the mass of NaCl(s) that results from the reaction of 55.4 g of $BaCl_2$ with an excess of Na_2SO_4 according to the following reaction.

$$Na_2SO_4(aq) + BaCl_2(aq) \rightarrow BaSO_4(s) + 2NaCl(aq)$$



Critical Item 4 - created on June 12, 20	800
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Calculate the mass of NaCl(s) that results from the reaction of 3.80 g of $BaCl_2$ with an excess of Na_2SO_4 according to the following reaction.

$$Na_2SO_4(aq) + BaCl_2(aq) \rightarrow BaSO_4(s) + 2NaCl(aq)$$



Critical Item 4 - created on June 12, 20	800
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Critical	item	4

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Calculate the mass of NaCl(s) that results from the reaction of 4.55 g of Na₂SO₄ with an excess of BaCl₂ according to the following reaction.

$$Na_2SO_4(aq) + BaCl_2(aq) \rightarrow BaSO_4(s) + 2NaCl(aq)$$

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Calculate the mass of NaCl(s) that results from the reaction of 780 g of Na_2SO_4 with an excess of $BaCl_2$ according to the following reaction.

$$Na_2SO_4(aq) + BaCl_2(aq) \rightarrow BaSO_4(s) + 2NaCl(aq)$$

Critical Item 4 - created or	n June 12, 2008
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Calculate the mass of CO_2 that results from the reaction of 780 g of O_2 with an excess of C_2H_5OH according to the following reaction.

$$C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O$$

Critical	Item 4	l - created	on J	June	12,	2008

Calculate the mass of H_2O that results from the reaction of 780 g of O_2 with an excess of C_2H_5OH according to the following reaction.

$$C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O$$

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Calculate the mass of H_2O that results from the reaction of 180 g of C_2H_5OH with an excess of O_2 according to the following reaction.

$$C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O$$

Critical Item 4 - created on June	12, 2008
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Calculate the mass of CO_2 that results from the reaction of 180 g of C_2H_5OH with an excess of O_2 according to the following reaction.

$$C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O$$