Critical Item #7

NAME _____

H_2SO_4 is titrated with NaOH according to the following reaction:

$2H_3PO_4 + 3Ca(OH)_2 \rightarrow Ca_3(PO_4)_2 + 6\ H_2O$

The concentration of the H_3PO_4 was 0.1252 M. 25.00 mL of H_3PO_4 is used and 31.52 mL of $Ca(OH)_2$ was required to reach the endpoint. What is the concentration of the $Ca(OH)_2$?

ANS _____M

Critical Item #7

NAME _____

H_2SO_4 is titrated with NaOH according to the following reaction:

$H_2SO_4 + 2NaOH \rightarrow Na_2SO_4 + 2\ H_2O$

The concentration of the NaOH was 0.1252 M. 25.00 mL of H_2SO_4 is used and 31.52 mL of NaOH was required to reach the endpoint. What is the concentration of the H_2SO_4?

ANS _____M

Critical Item #7

NAME _____

H_2SO_4 is titrated with NaOH according to the following reaction:

$H_2SO_4 + 2NaOH \rightarrow Na_2SO_4 + 2\ H_2O$

The concentration of the NaOH was 0.1252 M. 25.00 mL of NaOH is used and 11.67 mL of H_2SO_4 was required to reach the endpoint. What is the concentration of the H_2SO_4?

ANS _____M

Critical Item #7

NAME _____

H_2SO_4 is titrated with NaOH according to the following reaction:

$2H_3PO_4 + 3Ca(OH)_2 \rightarrow Ca_3(PO_4)_2 + 6\ H_2O$

The concentration of the $Ca(OH)_2$ was 0.1252 M. 25.00 mL of H_3PO_4 is used and 31.52 mL of $Ca(OH)_2$ was required to reach the endpoint. What is the concentration of the H_3PO_4?

ANS _____M

Critical Item #7

NAME _____

H_2SO_4 is titrated with NaOH according to the following reaction:

$2HCl + Ca(OH)_2 \rightarrow CaCl_2 + 2\ H_2O$

The concentration of the $Ca(OH)_2$ was 0.1252 M. 25.00 mL of HCl is used and 31.52 mL of $Ca(OH)_2$ was required to reach the endpoint. What is the concentration of the HCl?

ANS _____M

Critical Item #7

NAME _____

H_2SO_4 is titrated with NaOH according to the following reaction:

$2HCl + Ca(OH)_2 \rightarrow CaCl_2 + 2\ H_2O$

The concentration of the HCl was 0.1252 M. 25.00 mL of HCl is used and 31.52 mL of $Ca(OH)_2$ was required to reach the endpoint. What is the concentration of the $Ca(OH)_2$?

ANS _____M