Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

Calculate the pressure of 4.50 mol of an ideal gas which is contained in a 7.0 L volume at 200°C.

ANS: _____ atm

Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

Calculate the container volume if 3.00 mol of an ideal gas which is constrained with 6.0 atm pressure at 300°C.

ANS: _____ L

Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

Calculate the temperature of 3.00 mol of an ideal gas which is contained in a 5.0 L volume at a pressure of 15.0 atm. Show the temperature unit of your answer

ANS: _____ _____
unit !

Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

How many moles of an ideal gas which is contained in a 5.0 L volume at 270°C would generate 2.0 atm of pressure?

ANS: _____ mol

Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

Calculate the pressure of 2.50 mol of an ideal gas which is contained in a 17.0 L volume at 200°C.

ANS: _____ atm

Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

Calculate the container volume if 5.00 mol of an ideal gas which is constrained with 60.0 atm pressure at 300°C.

ANS: _____ L

Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

Calculate the temperature of 4.00 mol of an ideal gas which is contained in a 5.0 L volume at a pressure of 25.0 atm. Show the temperature unit of your answer

ANS: _____ _____
unit !

Critical Question #8

Note that 0 °C = +273 K and R = 0.08206 L atm mol^{-1} K^{-1}

Name _____

How many moles of an ideal gas which is contained in a 5.0 L volume at 250°C would generate 2.0 atm of pressure?

ANS: _____ mol