Calculate the pH to the nearest 0.01 unit of a 0.015 M solution of CH_3COOH. The K_a for CH_3COOH is 1.8 x 10^{-5}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 3.28

Calculate the pH to the nearest 0.01 unit of a 0.025 M solution of HCOOH. The K_a for HCOOH is 1.76 x 10^{-4}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 2.68

Calculate the pH to the nearest 0.01 unit of a 0.050 M solution of HNO_2. The K_a for HNO_2 is 4.0 x 10^{-4}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 2.35

Calculate the pH to the nearest 0.01 unit of a 0.35 M solution of $HClO_2$. The K_a for $HClO_2$ is 1.5 x 10^{-4}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 2.14

Calculate the pH to the nearest 0.01 unit of a 0.0025 M solution of C_6H_5COOH. The K_a for C_6H_5COOH is 6.30 x 10^{-5}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 3.40

Calculate the pH to the nearest 0.01 unit of a 0.075 M solution of HCOOH. The K_a for HCOOH is 1.8 x 10^{-4}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 2.43

Calculate the pH to the nearest 0.01 unit of a 0.0085 M solution of CH_3COOH. The K_a for CH_3COOH is 1.8 x 10^{-5}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 3.41

Calculate the pH to the nearest 0.01 unit of a 0.15 M solution of HIO_2. The K_a for HIO_2 is 1.8 x 10^{-6}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 3.28

Calculate the pH to the nearest 0.01 unit of a 0.0085 M solution of HClO. The K_a for HClO is 3.7 x 10^{-8}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 4.75

Calculate the pH to the nearest 0.01 unit of a 0.090 M solution of $HBrO_2$. The K_a for $HBrO_2$ is 3.8 x 10^{-7}. Write the Brønsted-Lowery reaction and the equilibrium expression.
Ans = 3.73