Critical Item 7 – General Chemistry II copy 1

NAME_____

Calculate the molar solubility of Cu_2S. $K_{sp} = 2.0 \times 10^{-47}$.
1. Give the right number to 2 significant figures.

Write the equilibrium reaction here:_____

Show you calculation here clearly

pH = _____

Critical Item 7 – General Chemistry II copy 2

NAME_____

Calculate the molar solubility of $Fe(OH)_2$. $K_{sp} = 9.0 \times 10^{-35}$.
Give the right number to 2 significant figures.

Write the equilibrium reaction here:_____

Show you calculation here clearly

pH = _____

Critical Item 7 – General Chemistry II copy 3

NAME_____

Calculate the molar solubility of $Sn(OH)_4$. $K_{sp} = 1.0 \times 10^{-56}$.
Give the right number to 2 significant figures.

Write the equilibrium reaction here:_____

Show you calculation here clearly

pH = _____

KEY

copy 1 $[Cu_2S] = 1.7 \times 10^{-16}$
copy 2 $[Fe(OH)_2] = 2.8 \times 10^{-12}$
copy 3 $[Sn(OH)_4] = 2.1 \times 10^{-12}$