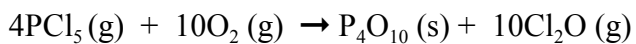
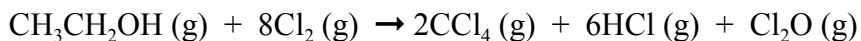


Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



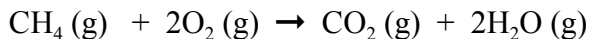
$$\text{Ans} = -497 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



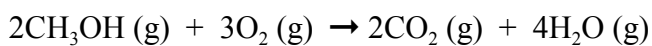
$$\text{Ans} = -426.3 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298 K for the following reaction:



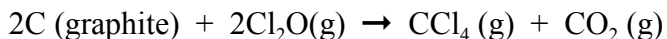
$$\text{Ans} = -800.8 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



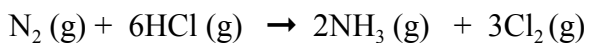
$$\text{Ans} = -1379.18 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



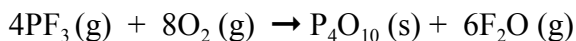
$$\text{Ans} = -650.94 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



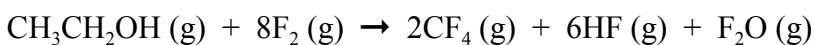
$$\text{Ans} = +538.7 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



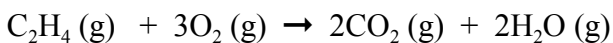
$$\text{Ans} = +864.27 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



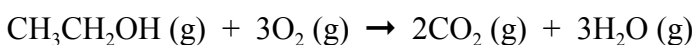
$$\text{Ans} = -3235.3 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



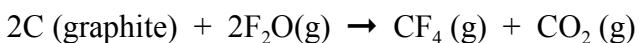
$$\text{Ans} = -1313.97 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



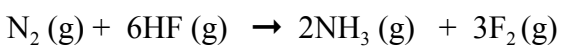
$$\text{Ans} = -1305.8 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



$$\text{Ans} = -1265.4 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



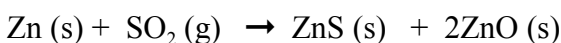
$$\text{Ans} = +1606.14 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



$$\text{Ans} = -80.30 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:



$$\text{Ans} = -562.7 \text{ kJ mol}^{-1}$$

Calculate the  $\Delta G^\circ$  at 298K for the following reaction:

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$$\text{Ans} = +471.5 \text{ kJ mol}^{-1}$$